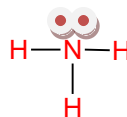


1. Write Lewis dot structures for the following molecules: (4 pts)

a) H₂



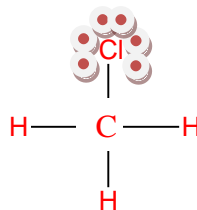
b) NH₃



c) N₂



d) CH₃Cl



2. Classify each of the following bonds as ionic, covalent, or polar covalent. (4 pts)

a. H-F

$$4.0 - 2.1 = 1.9$$

Polar Covalent

b. C-N

$$3.0 - 2.5 = 0.5$$

Polar Covalent

Electronegativity

Atom	Value
C	2.5
H	2.1
F	4.0
O	3.5
N	3.0
Cl	3.0
Li	1.0

c. Li-Cl

$$3.0 - 1.0 = 2.0$$

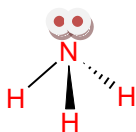
Ionic

d. O-O

$$3.5 - 3.5 = 0$$

Covalent

3. Indicate both the electron geometry and molecular geometry of NH₃. (2 pts)



Electron Geometry = Tetrahedral

Molecular Geometry = Trigonal Pyramidal