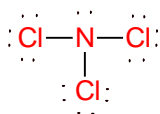
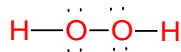


1. Write Lewis structures for the following: (4 pts)

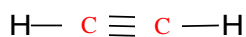
a) NCl_3



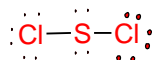
b) H_2O_2



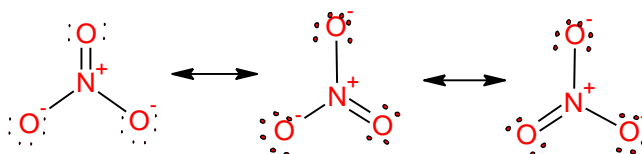
c) C_2H_2



d) SCl_2



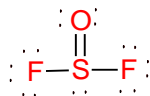
2. Write resonance structures for NO_3^- and show the formal charges on the atoms. (3 pts)



3 stable resonance structures

3. For SOF_2 , write the most stable Lewis structure. (3 pts)

The most stable structure would not have any formal charges on the atoms.



Sulfur may violate the octet rule because it has an empty set of 3d orbital that may accept electrons.