

1. Fill in the missing information in the following table: (4pts)

Symbol	Protons	Neutrons	Electrons	Charge
^{85}Rb	37	48	37	0
$^{32}\text{S}^{2-}$	16	16	18	-2
$^{40}\text{Ca}^{2+}$	20	20	18	+2
$^{51}\text{V}^{+n}$	23	28	19	+4

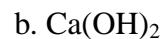
2. An element has three naturally occurring isotopes shown below:

Isotopic Mass (amu)	Fractional Abundance
28	0.9221
29	0.0470
30	0.0309

What is the atomic weight of this element? (2 pts)

$$\text{Atomic weight} = (28)(0.9221) + (29)(0.0470) + (30)(0.0309) = 28.1088 \text{ Amu}$$

3. Name the following compounds:(2 pts)



Ammonium Perchlorate

Calcium Hydroxide

4. Which of the following elements are
- isoelectronic
- to krypton? (2 pts)

