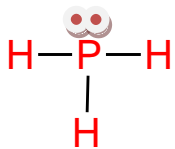
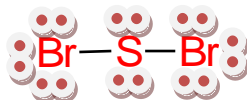


1. Write Lewis structure for the following: (4 pts)

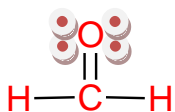
a) PH<sub>3</sub>



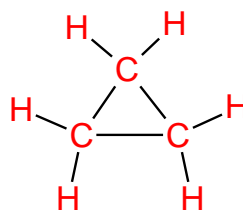
b) SBr<sub>2</sub>



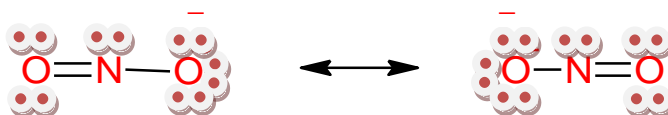
c) CH<sub>2</sub>O



d) C<sub>3</sub>H<sub>6</sub>



2. Write the resonance structures for NO<sub>2</sub><sup>-</sup> and show the formal charges on all the charged atoms. (3 pts)



3. Draw the most stable Lewis structure for the following molecule and show all formal charges on the atoms. (3 pts)

PO<sub>4</sub><sup>3-</sup>

