Name ____Mr. Perfect_____ Date _____F11_____

1) Fill in the missing information in the following table: (4pts)

Symbol	Protons	Neutrons	Electrons	Charge
^{9}Be	4	5	4	0
$^{19}F^{-}$	9	10	10	-1
$^{39}K^{+}$	19	20	18	+1
$^{55}Mn^{+n}$	25	30	20	+5

2) An element has two naturally occurring isotopes with the following masses and abundances:

Isotopic Mass (amu)	Fractional Abundance
106.905	0.5184
108.905	0.4816

What is the atomic weight of this element? (2 pts)

Atomic weight = (0.5184)(106.905) + (0.4816)(108.905) = **107.869 Amu**

3) Name the following compounds:(2 pts)

a. $CoCl_3$ b. BrO_3^-

Cobalt(III) Chloride Bromate

4) Using the "Bohr Model" of the atom, draw the electron dot structure for the following element. (2 pts)

^{24}Mg atomic number = 12

12 protons
12 electrons