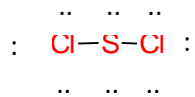
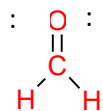


1. Write the Lewis structures for the following compounds: (4 pts)

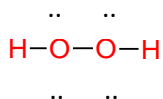
a. SCl_2



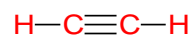
b. CH_2O



c. H_2O_2

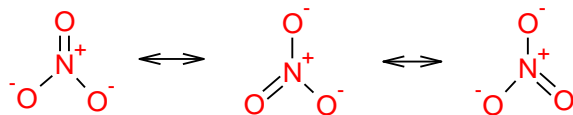


d. C_2H_2



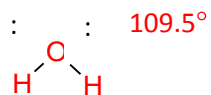
2. Write the stable resonance structures for NO_3^- polyatomic ion and show the formal charges on the atoms. (3 pts)

3 resonance structures (lone pair electrons not shown)



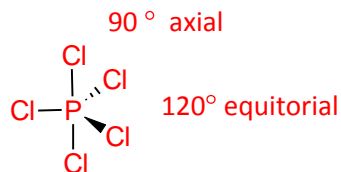
3. Using VSEPR theory, draw the following in the correct molecular geometry and include the bond angles. Also, give the name of the molecular geometry. (3 pts)

a. H_2O



Bent

b. PCl_5



Trigonal Bipyramidal