

065 Extra Credit Practice Exam

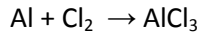
12. The formula for a molecule formed from N and Cl would be:

- a. NCl b. NCl₂ c. NCl₃ d. N₃Cl

13. What is the formula for aluminum nitrate?

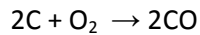
- a. Al₂NO₂ b. AlNO₃ c. Al(NO₂)₃ d. Al(NO₃)₃

14. In the following reaction, what is the correct coefficient for aluminum chloride?



- a. 1 b. 2 c. 3 d. 4

15. What is the classification for this reaction?



- a. single replacement b. double replacement c. combination d. decomposition

16. How many moles of carbon are there in 0.500 moles of C₂H₆?

- a. 0.500 moles b. 1.00 moles c. 3.00 moles d. 4.00 moles

17. The pressure of 5.0 L of gas increases from 1.50 atm to 1249 mmHg. What is the final volume of the gas, assuming constant temperature?

- a. 4100L b. 5.0 L c. 5.4 L d. 4.6 L

18. How many moles of neon occupy a volume of 14.3 L at STP?

- a. 36.7 moles b. 32.0 moles c. 6.45 moles d. 0.638 moles

19. What is the molarity of a solution that contains 17 g of NH₃ in 0.50 L of solution?

- a. 34 M b. 2.0 M c. 0.50 M d. 0.029 M

20. What is the (m/v)% of KOH when 5.78 g of KOH are dissolved into 250 mL of water?

- a. 7.3 % b. 2.3 % c. 5.0 % d. 1.6 %

***Bring a scantron and a number 2 pencil to the extra credit project. The extra credit exam is 50 questions long.*